

The effect of the linker size in C₂-symmetrical chiral ligands on the self-assembly formation of luminescent triple-stranded di-metallic Eu(III) helicates in solution†

Oxana Kotova^{*a}, Steve Comby^a, Komala Pandurangan^a, Floriana Stomeo^a, John E. O'Brien^a, Martin Feeney^a, Robert D. Peacock^b, Colin P. McCoy^c and Thorfinnur Gunnlaugsson^{*a}

^aSchool of Chemistry, Trinity Biomedical Sciences Institute (TBSI), Trinity College Dublin, Dublin 2, Ireland.

^bSchool of Chemistry, University of Glasgow, Glasgow, G 12 8QQ Scotland, UK

^cSchool of Pharmacy, Queen's University Belfast, 97 Lisburn Road, Belfast, BT9 7BL Northern Ireland, UK

E-mail: kotovao@tcd.ie, gunnlaut@tcd.ie

Abstract

Chiral lanthanide-based supramolecular structures have gained significant importance in view of their application in imaging, sensing and other functional purposes. We have designed chiral C₂-symmetrical ligands (L) based on the use of two 2,6-pyridine-dicarboxylic-amide moieties (pda), that differ from one another by the nature of the diamine spacer groups (from 1,3-phenylenedimethanamine (1(S,S), 2(R,R)) and benzene-1,3-diamine (3(S,S), 4(R,R)) to much bulkier 4,4'-(cyclohexane-1,1-diyl)bis(2,6-dimethyl-aniline) (5(S,S), 6(R,R))) between these two pda units. The self-assembly between L and Eu(III) ions were investigated in CH₃CN solution at low concentration whereby the changes in the absorbance, fluorescence and Eu(III)-centred emission spectra allowed us to model the binding equilibria occurring in

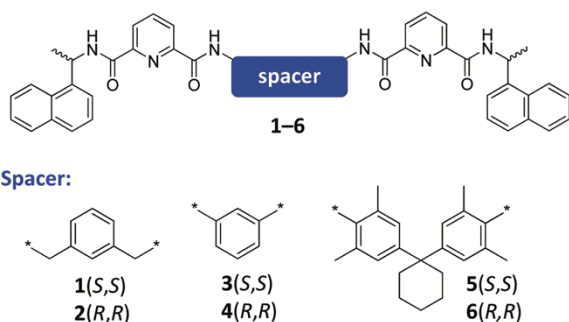
the solution to the presence of [Eu:L2], [Eu₂:L2], [Eu₂:L3] assemblies and reveal their high binding constant values. The self-assembly in solution were also studied at higher concentration by following the changes in the ¹H NMR spectra of the ligands upon Eu(III) addition, as well as by using MALDI-MS of the isolated solid state complexes. The chiroptical properties of the ligands were used in order to study the structural changes upon self-assembly between the ligands and Eu(III) ions using circular dichroism (CD) and circularly polarised luminescence (CPL) spectroscopies. The photophysical properties of [Eu₂:L3] complexes were evaluated in solution and showed a decrease of luminescence quantum yield when going from the ligand with smaller (1(S,S)) to bulkier (5(S,S)) linker from ~5.8% to ~2.6%. While mass-spectrometry revealed the possible formation of trinuclear assemblies such as [Eu₃:L3] and [Eu₃:L2].

Introduction

The design of self-assembled lanthanide (Ln) complexes with triple-stranded di-metallic helical motif is inspired by the architecture, mechanics and structural features of biological macromolecules and driven by the possible application of such systems thanks to the unique magnetic and spectroscopic properties of the metal centres.¹ The first di-metallic lanthanide helicates were reported by Piguet et al.² and since then, the structural and thermodynamic parameters of such systems were investigated³ moving towards nd-4f helicates,⁴ polymeric helical structures,⁵ and water soluble systems for imaging applications.⁶ Such systems were also developed as sensors for adenosine monophosphate, fluoride⁷ and chiral helical structures; which we also investigate within our research group.⁸ The design of the ligands is crucial in order to achieve lanthanide complexes with triple-stranded di-metallic helical motif and modifications of ligand design have been recently proven successful in moving from triple stranded helicates towards achieving 3D helical structures,⁹ pentanuclear lanthanide

helicates,¹⁰ tetranuclear tetrahedrons and octanuclear cubes,¹¹ or cages.¹² Recently, another interesting approach has been shown where one can move from the use of dinuclear bis- β -diketonate lanthanide complexes,¹³ to D₂-symmetrical alternating circular helicate, by simply combining achiral bis- β -diketonate and chiral Pybox ligands within a single ligand structure.¹⁴ In our research group, we have based our ligand design on connecting two chiral half-helicate ligands either (S) or (R) 6-(1-(naphthalen-1-yl)ethylcarbamoyl)picolinic acid with diamine spacers such as 4,4'-methylenedianiline,^{8a} 3,3'-methyl-enedianiline^{8b} or smaller 1,3-phenylenedi-methanamine.^{8c} The interaction of these ligands with lanthanide ions in solution resulted in the formation of enantiomerically pure luminescent triple-stranded di-metallic Ln(III) helicates. The use of 2,6-pyridinedicarboxamide (pda) binding unit is very appealing for the formation of various types of Ln(III) bundles^{8g} including highly novel systems such as molecular trefoil knots,^{15a,b} elegantly developed by Leigh and co-workers, that can be used in asymmetric catalysis.^{15c} This concept was recently expanded for Zn(II) systems where trefoil knots were formed from trimeric circular helicates,^{15d} and by securing a supramolecular architecture by tying a stopper in mechanically interlocked molecules.^{15e} In this work we continue our helicate endeavour with Eu(III) and investigate the effect of the spacer size on the formation of lanthanide self-assembly structures in solution by comparing the system published earlier (ligands 1–2)^{8c} with the new ligands where shorter benzene-1,3-diamine (ligands 3–4, Scheme 1) or much bulkier spacer such as 4,4'-(cyclohexane-1,1-diyl)bis(2,6-dimethylaniline) (ligands 5–6) were utilised (Scheme 1).

The structural changes of the ligand upon self-assembly with Eu(III) in solution were accessed with the aid of spectroscopic techniques by monitoring the changes in the absorbance, fluorescence and Eu(III)-centred emission of the ligand upon metal addition at low concentration, as well as circular dichroism (CD) and circularly polarised luminescence (CPL) spectroscopies, in addition to ¹H NMR spectroscopy at higher ligand concentrations.



Scheme 1

Structural formulas of the ligands 1–6.

Results and discussion

Synthesis of ligands 1–6

The ligands were synthesised in four steps (Scheme S1, ESI†) starting from 2,6-pyridine-dicarboxylic acid through protecting one of the carboxylic acids by reacting it with benzyl bromide. The remaining carboxylic acid group, of resulting compound 16 (see ESI, Scheme S1†) was coupled with (*S*) or (*R*) 1-(naphthalen-1-yl)ethanamine using *N*-(3-dimethylaminopropyl)-*N'*-ethylcarbodiimide (EDCI), 1-hydroxybenzotriazole (HOBt) as coupling reagents in dichloromethane resulting in the formation of compounds 12, 13 (Scheme S1, ESI†), followed by de-protection using hydrogenation in the presence of or Pd/C as a catalyst. These so-called “half-helicate” compounds (either 10 or 11, Scheme S1, ESI†) were then coupled with the desired “spacer” diamines (Scheme 1) using the same peptide coupling reagents as before, in the presence of trimethylamine and 4-dimethylaminopyridine as catalyst. It has to be noted that 1,3-phenylenediamine linker (8) was obtained from commercial sources while 4,4'-(cyclohexane-1,1-diyl)bis(2,6-dimethylaniline) linker (9) was synthesised following the procedures reported previously by Hunter et al.¹⁷ Ligands 3–6 (Scheme 1) were obtained in ~20–25% yield and their composition was confirmed by a

combination of ^1H , ^{13}C and heteronuclear single quantum correlation (HSQC) NMR experiments along with electrospray mass-spectrometry (ESI-MS), IR and elemental analysis (see ESI, Fig. S1–S14[†]).

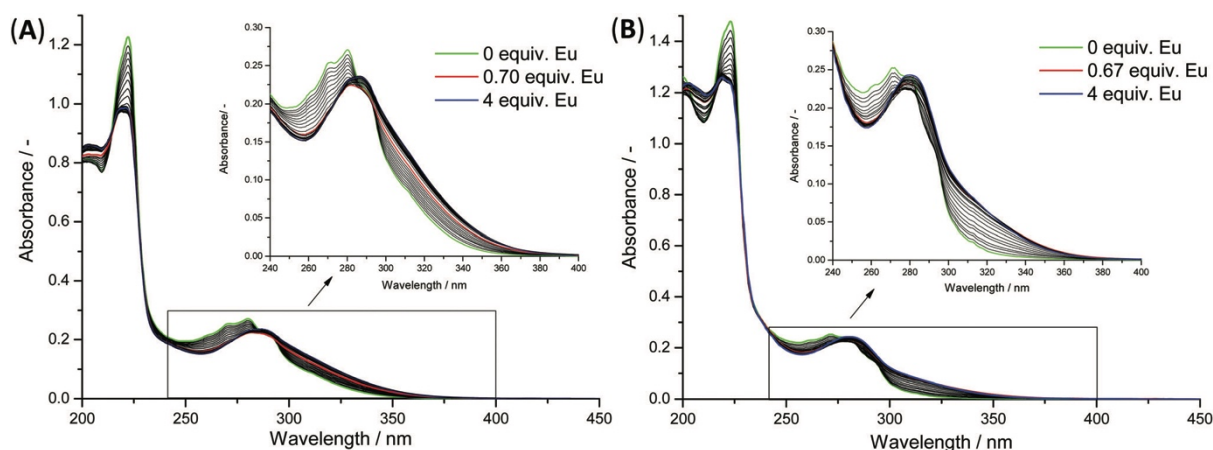


Fig. 1 The changes in the absorption spectrum of ligands (A) 4(R,R) ($c = 7.84 \times 10^{-6}$ M) or (B) 5(S,S) ($c = 1.00 \times 10^{-5}$ M) upon addition of $\text{Eu}(\text{OTf})_3$ (0 \rightarrow 4 equiv.) in CH_3CN at 25 $^\circ\text{C}$.

Self-assembly studies between chiral ligands 3–6 and $\text{Eu}(\text{III})$ ions in solution, formation of helical assemblies at low concentration using UV-Vis and luminescence spectroscopy

Initially, the absorption spectra of the ligands developed herein, were evaluated in CH_3CN solution (Fig. S15[†]). The same trend was seen for all, where two main absorption bands centred at ca. 223 nm and 280 nm were observed, these being assigned to $\pi \rightarrow \pi^*$ and $n \rightarrow \pi^*$ transitions within ligands structure ($\epsilon_{280}(3(\text{S,S})) = 34553 \text{ M}^{-1} \text{ cm}^{-1}$, $\epsilon_{280}(4(\text{R,R})) = 34963 \text{ M}^{-1} \text{ cm}^{-1}$, $\epsilon_{280}(5(\text{S,S})) = 23260 \text{ M}^{-1} \text{ cm}^{-1}$ and $\epsilon_{280}(6(\text{R,R})) = 23358 \text{ M}^{-1} \text{ cm}^{-1}$).

The self-assembly studies between the ligands 3–6 and $\text{Eu}(\text{III})$ ions in CH_3CN solution were next monitored by following the changes in the absorption, fluorescence and $\text{Eu}(\text{III})$ -centred

emission spectra of each ligand upon gradual addition of metal ions ($0 \rightarrow 4$ equivalents) (Fig. 1). The changes in the titration profiles for the pair of enantiomers 3(S,S), 4(R,R) and 5(S,S), 6(R,R) followed the same trends, and therefore, we will only discuss one of each sets of enantiomers herein. The changes in the absorption spectra of 4(R,R) and 5(S,S) are shown in Fig. 1A and B; the long absorption wavelength being red shifted upon complexation of these ligands with Eu(III). The changes observed for 3(S,S) and 6(R,R) upon addition of Eu(III) ($0 \rightarrow 4$ equiv.) were to greater extend identical; the main differences being the observation of a less decrease in the absorbance centred at 222 nm between the additions of $0 \rightarrow 0.50$ equivalents of Eu(III), and higher hyperchromicity of the band centred at 334 nm in the case of 5(S,S) and 6(R,R) compared to 3(S,S) and 4(R,R) (Fig. 1A and B). This can reflect the differences in the conformational changes between the self-assembly processes in solution. In both cases the changes in the absorption bands of the ligand occurred until addition of 0.67 equivalents of Eu(III) when the plateau has been reached (Fig. S16, S17, S21 and S24†). Excitation into the naphthalene bands at $\lambda_{\text{ex}} = 281$ nm resulted in a weak ligand-centred emission with the maximum at $\sim 425\text{--}460$ nm for all the ligands (Fig. S18, S19, S22 and S25†). Upon addition of Eu(III) ($0 \rightarrow 0.67$ equivalents) an enhancement was seen in the ligand-centred emission of the benzene-1,3-diamine linker based ligands 3(S,S) and 4(R,R), while after the addition of 0.67 equivalents of Eu(III), a gradual decrease in the ligand-centred emission was observed. Concomitantly, the Eu(III)-centred emission was observed, indicating the successful population of the lanthanide excited state, through sensitisation of the naphthalene antennae, with the maximum of luminescence intensity being achieved upon the addition of ca. 0.70 equivalents of Eu(III). Upon further additions, a decrease in the lanthanide luminescence was observed until it reached plateau after the addition of ca. 1.25 equivalents (Fig. S18 and S22). The case of increasing ligand centred emission upon complexation with Eu(III) is unusual and suggests the presence of strong interaction between

the ligands as well as decrease of vibration quenching effects within ligand structure occurring upon complexation.¹⁶ In the case of 5(S,S) and 6(R,R) (with 4,4'-(cyclohexane-1,1-diyl)bis(2,6-dimethylaniline) linker)¹⁷ the weak ligand centred fluorescence decreased upon complexation with Eu(III). At the same time, the Eu(III)-centred emission appears; reaching maximum upon addition of 0.67 equivalents of Eu(III). After this point, the ligand-centred emission is gradually enhanced, while Eu(III)-centred emission decreased until 4 equivalents of the ion added (Fig. S19 and S25†). The changes in the delayed (0.10 ms) Eu(III)-centred emission were following the same trend as these observed in the fluorescence spectra earlier (Fig. 2, S20, S23 and S26†). The ratio between the 5D0 → 5FJ transitions for these two sets of diamine linkers varies significantly, suggesting discrepancy in the coordination environment of Eu(III) ion. Interestingly, the ratio between the species formed in solution also varies, which is a sign that different self-assembly processes are occurring in solution for these two types of ligands. However, in both cases, the presence of 5D0 → 7F0 band is observed, this suggesting the formation of species where the Eu(III) ions are within a site symmetry of C_{nv}, C_n or C_s.¹⁸ In contrast to these titrations, the titration profile for 3(S,S) and 4(R,R) (with benzene-1,3-diamine linker) upon titration with Eu(III) corresponds closer to the behaviour previously observed for 1(S,S) and 2(R,R) (with 1,3-phenylenedimethanamine).^{8c} In all the cases the shape of the excitation spectra with λ_{max} at 280 nm recorded during the course of the titrations (λ_{em} = 616 nm; Fig. S27†) confirmed that the population of the Eu(III) excited state occurs via the “antenna effect” where the energy absorbed by the ligands is transferred onto the metal centres from where the characteristic sharp line-like emission occurs. During the course of the titration the lifetimes of the 5D0 level of Eu(III) were recorded upon the addition of 0.67 eq. and these results were then best fitted to mono-exponential decays with the average values of ~1.50–1.70 ms suggesting the formation of fully saturated metal coordination spheres (Fig. S28†).^{8a,b,19a,b}

The changes observed during the course of these titrations were analysed by fitting the data using non-linear regression analysis program SPECFIT.²⁰ In all the cases, the factor analysis suggested the formation of three additional species along with the initial ligand being present in solution. The stability constants were determined (as $\log \beta_{ML}$) for the binding models [Eu:L2], [Eu₂:L2] and [Eu₂:L3] (Table 1). The values of the binding constants between the enantiomers of the same ligand closely corresponded to each other, and were found to be higher for the ligands consisting of smaller linker groups such as (1(S,S), 2(R,R))**8c** and (3(S,S), 4(R,R)) compared to (5(S,S), 6(R,R)) (Table 1). The speciation distribution diagram for these titrations also revealed that the yield of [Eu₂:L3] species at 0.67 equiv. of Eu(III) decreased from 80.2% for 1(S,S), 2(R,R)**8c** and 82.5% for 3(S,S), 4(R,R) to 72.2% for 5(S,S), 6(R,R), Fig. S29.[†]

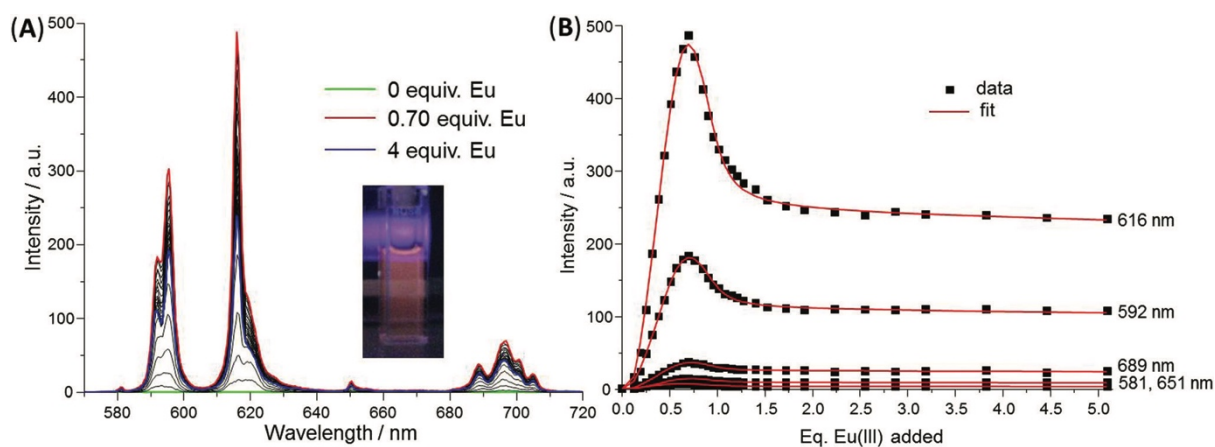


Fig. 2 (A) The changes in the delayed Eu(III)-centred emission spectrum of 4(R,R) ($c = 7.84 \times 10^{-6}$ M) upon addition of Eu(OTf)₃ (0 \rightarrow 4 equiv.) in CH₃CN at T = 25 °C, $\lambda_{ex} = 281$ nm (total decay time = 0.02 s; delay time = 0.1 ms); (B) experimental binding isotherms for the same titration and their corresponding fits by means of SPECFIT program.

Circular dichroism (CD) and circularly polarised luminescence (CPL) for [Eu₂:L₃] assemblies in solution

The chiral nature of the ligands was probed using CD spectroscopy, and the Cotton effects in the spectra of corresponding pair of enantiomers (e.g. 3(S,S), 4(R,R) and 5(S,S), 6(R,R)) appeared as mirror images to each other, Fig. 3. For 3(S,S), three positive CD signals were observed with the maxima at ca. 296, 229, 208 nm along with two negative CD signals at 271 and 219 nm. The signals in the CD spectra of 5(S,S) were similar to these observed for 3(S,S) with the positive signals observed at 285, 230 and 209 nm, correspondingly, while showing only one negative signal at 223 nm. The shape of the CD spectra observed here closely corresponds to similar structures previously developed.^{8a,b}

The corresponding [Eu₂:L₃] were then formed in solution through self-assembly of 1 equiv. of the corresponding ligand with 0.67 equiv. of Eu(III). Each solution was then equilibrated for 15 minutes to ensure the formation of the desired [Eu₂:L₃] helical assemblies in solution (Fig. 3). Again the spectra corresponding to each pair of enantiomers appeared as mirror image to one another suggesting enantiomeric purity of the corresponding Eu(III) species (although this might also represent a minor contribution from other stoichiometries as outlined above). The CD spectra of the complexes appeared significantly different to those of the ligands. For example, for [Eu₂:L₃] with L = 3(S,S) (Fig. 3A) the first negative Cotton effect band was observed at 325 nm while second positive Cotton effect band at 295 nm and an amplitude of 4.6 mdeg showing negative chirality. The following positive Cotton effect centred at 227 nm and negative Cotton effect at 214 appeared with large amplitude of 115 mdeg. The presence of sinusoid-like shape of the CD spectra for [Eu₂:L₃] suggests the appearance of possible mixing between electric- and magnetic-dipole allowed transitions (most likely exciton coupling) between the chromophores that has been brought in one another close proximity due to the formation of the Eu(III) assemblies.²¹ Similarly, for

[Eu2:L3] complexes where L = 5(S,S) (Fig. 3B), the first negative Cotton effect was observed at 340 nm, with the following positive Cotton effect being observed at 295 nm and the amplitude of 3 mdeg. The latter positive Cotton effect then evolved into the next positive CD band centred at 227 nm with the following negative Cotton effect at 212 nm and the amplitude of 245 mdeg. Similarities in the shape of the CD spectra between [Eu2:L3] (L = 3(S,S), 4(R,R); Fig. 3A) and (L = 5(S,S), 6(R,R); Fig. 3B) suggests similar geometrical character of the assemblies being formed in CH3CN.

Table 1 Binding constants ($\log \beta_{M:L}$) obtained by fitting the changes in absorption, fluorescence and Eu(III)-centred emission spectra of 3(S,S), 4(R,R), 5(S,S) or 6(R,R) upon addition of Eu(CF3SO3)3 in CH3CN solution (25 °C). All titrations were done at least in triplicates

Eu(III)	Absorbance			Fluorescence			Eu(III)-Centred emission		
	$\log \beta_{1:2}$	$\log \beta_{2:3}$	$\log \beta_{2:2}$	$\log \beta_{1:2}$	$\log \beta_{2:3}$	$\log \beta_{2:2}$	$\log \beta_{1:2}$	$\log \beta_{2:3}$	$\log \beta_{2:2}$
L = 3(S,S)	13.7 ^a	26.7 ± 0.2	19.6 ± 0.2	13.3 ± 0.4	26.6 ± 0.5	19.6 ± 0.4	13.2 ± 0.4	27.9 ± 0.2	21.2 ± 0.1
L = 4(R,R)	12.5 ± 0.4	26.4 ± 0.3	19.1 ± 0.2	13.2 ^a	27.1 ± 0.2	20.7 ± 0.1	13.2 ± 0.1	26.6 ± 0.2	20.7 ± 0.1
L = 5(S,S)	12.7 ± 0.1	25.5 ± 0.3	18.1 ± 0.2	14.3 ± 0.3	26.2 ± 0.3	19.2 ± 0.2	13.9 ± 0.2	26.2 ± 0.4	19.1 ± 0.1
L = 6(R,R)	12.9 ± 0.1	25.4 ± 0.3	19.3 ± 0.2	15.0 ± 0.2	28.5 ± 0.3	20.3 ± 0.3	13.9 ± 0.2	25.8 ± 0.6	19.9 ± 0.3

^a Values were fixed to allow convergence of the experimental data to the proposed model.

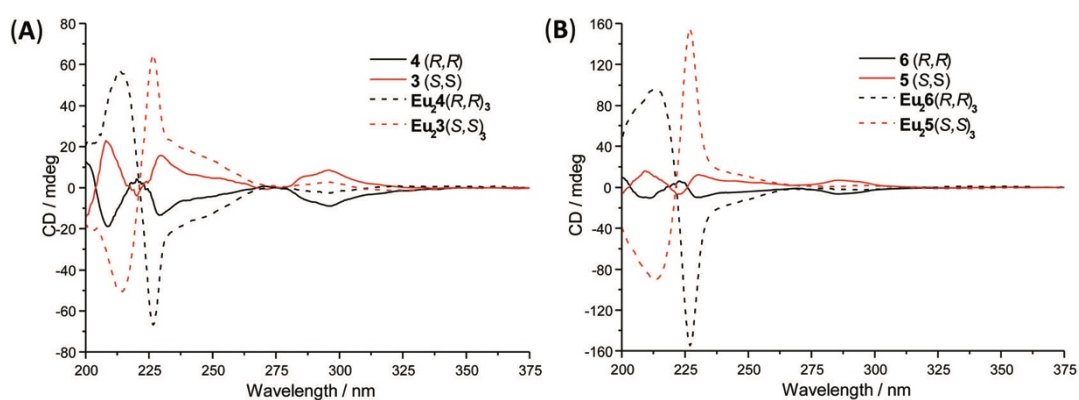


Fig. 3 Circular dichroism spectra of the ligands (A) L = 3(S,S), 4(R,R) ($c = 1.43 \times 10^{-5}$ M) and (B) L = 5(S,S), 6(R,R) ($c = 1.40 \times 10^{-5}$ M) together with their corresponding Eu₂L₃ complexes ($c_{\text{Eu}_2\text{L}_3} = 4.77 \times 10^{-6}$ M) in CH₃CN.

The changes in the CD spectra can be used to correlate the nonlinear enhancement of the chiroptical response to the non-linearity dependence on point chirality for tetrahedral Eu(III) chiral cage structures or used to monitor anion sensing events for different anions in solution.^{7b,12c} Furthermore, in our previous work,²² we have shown that the changes in the CD spectra of chiral ligands titrated with Ln(III) ions can be used not only to monitor the structural/conformational changes occurring upon complexation/self-assembly formation, but also used to evaluate the binding affinities between the ligands and the ions in solution by fitting the data using non-linear regression analysis. With this in mind, we monitored the changes in the CD spectra of the ligands L = 5(S,S) and 6(R,R), upon addition of Eu(III) (0 → 4 equiv.). From these titrations, Fig. 4, it was clear that a significant structural conformations occurred upon formation of the helicates in solution, and the presence of a kinetic effects was also observed.²³ As previously observed for CD spectra of the ligands and their assemblies with Eu(III) (Fig. 3) the titrations profile for 5(S,S) was a mirror image to that observed for 6(R,R) (Fig. 4), and the binding isotherms (Fig. S30†) reveal the main changes in the spectra happening until the addition of 0.67 equiv. of Eu(III) (Table 1), supporting the formation of the bi-metallic triple-stranded helicates in solution. Finally, the CPL spectra of [Eu₂:L₃] assemblies were recorded in CH₃CN solution, where the Eu(III) centred emission was monitored upon ligand excitation (Fig. 5). This clearly revealed the chiral nature of the Eu(III) centres; each pair of enantiomers resulted in CPL

spectra that were mirror images opposite in signs and equal in magnitudes. The absolute configuration of Eu(III) centres can be tentatively assigned by comparing these results to the CPL spectra of previously published mononuclear Eu(III) complexes, where the absolute configuration was determined by using X-ray crystallography.¹⁹ As such, the [Eu₂:L₃] self-assemblies where L = 3(S,S), 5(S,S) can be assigned as being Δ,Δ around the two metal ions, while for L = 4(R,R), 6(R,R), the assignment of the absolute configuration is Λ,Λ. The dissymmetry factors (glum) are listed in Table 2 and represented in the g-value plots on Fig. S31.† The 5D₀ → 7F₁ transition of these were determined to be of similar values observed for other reported mono-metallic complexes and di-metallic helicates.^{19a,b,12c} The exception, however, being the bulkier ligands 5(S,S) and 6(R,R) which gave lower glum values. The glum values determined for the 5D₀ → 7F₂ transitions were also found to be similar to the mono-metallic Eu(III) “Trinity Sliotar” bundles, previously developed in our group.^{19a} In comparison to these results, the L = 3(S,S), 4(R,R) assemblies showed relatively high glum values for the 5D₀ → 7F₃ transition; being 0.40–0.60, which is among the highest reported for this kind of structures. Additionally reasonably high glum values were observed for the 5D₀ → 7F₄ transition; being 0.45–0.53 for [Eu₂:L₃] (L = 3(S,S), 4(R,R)) and ~0.45 for [Eu₂:L₃] (L = 5(S,S), 6(R,R)) (Table 2). Such relatively high glum(D₀ → F₃) values might be a sign of strong J-mixing and a strong crystal-field perturbation in these triple-stranded helicates.^{18b}

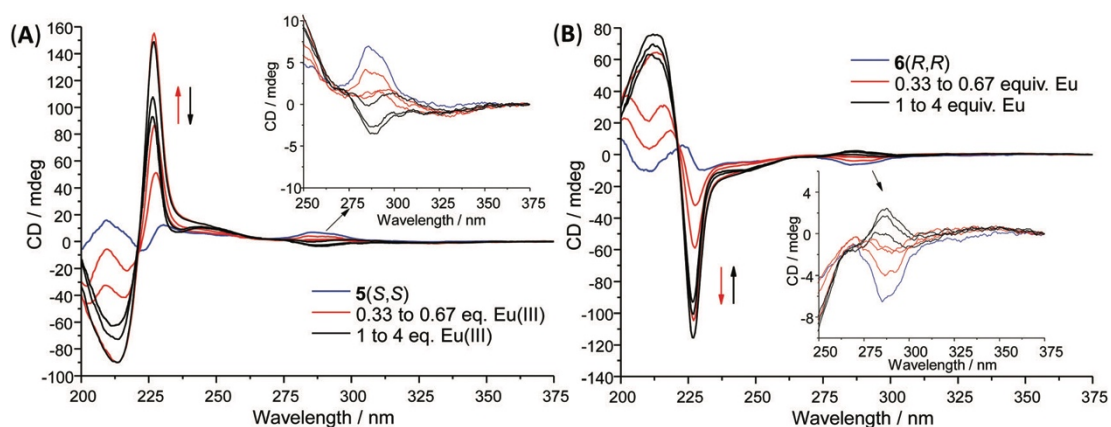


Fig. 4 The changes in the CD spectra of (A) 5(S,S) ($c = 1.40 \times 10^{-5}$ M) and (B) 6(R,R) ($c = 1.34 \times 10^{-5}$ M) upon addition of Eu(OTf)₃ (0 → 4 equiv.) in CH₃CN.

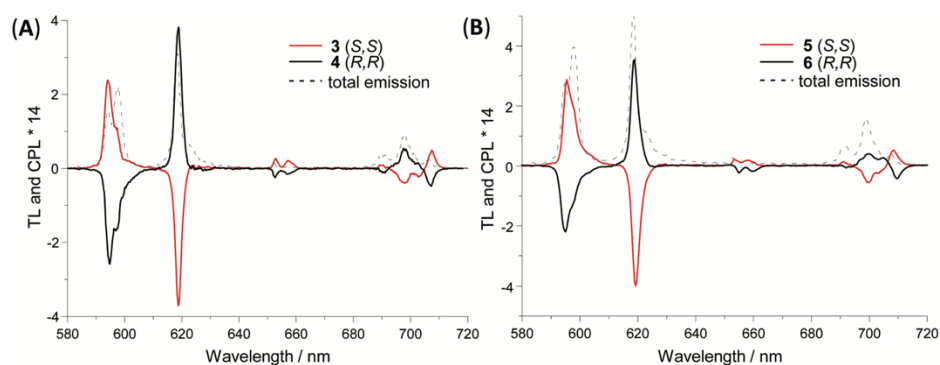


Fig. 5 Circularly polarised luminescence spectra and total luminescence emission of Eu₂L₃ complexes in CH₃CN ($c = 3.33 \times 10^{-5}$ M) where L is (A) 3(S,S), 4(R,R) and (B) L = 5(S,S), 6(R,R) ($\lambda_{\text{ex}} = 281$ nm).

Table 2 Dissymmetry factor values (g_{lum}) determined from the spectra depicted on Figure 3

Transition	λ , nm	$g_{\text{lum}}(\text{Eu}_2(3(S,S))_3)$	$g_{\text{lum}}(\text{Eu}_2(4(R,R))_3)$	$g_{\text{lum}}(\text{Eu}_2(5(S,S))_3)$	$g_{\text{lum}}(\text{Eu}_2(6(R,R))_3)$
$^5D_0 \rightarrow ^7F_1$	595	0.27	-0.27	0.17	-0.17
$^5D_0 \rightarrow ^7F_1$	603	0.19	-0.27	0.18	-0.16
$^5D_0 \rightarrow ^7F_2$	620	-0.27	0.28	-0.21	0.20
$^5D_0 \rightarrow ^7F_3$	654	0.43	-0.50	0.20	-0.26
$^5D_0 \rightarrow ^7F_3$	658	0.57	-0.61	0.32	-0.32
$^5D_0 \rightarrow ^7F_4$	696	-0.06	0.06	-0.03	0.05
$^5D_0 \rightarrow ^7F_4$	700	-0.08	0.09	-0.06	0.05
$^5D_0 \rightarrow ^7F_4$	704	-0.08	0.09	-0.07	0.10
$^5D_0 \rightarrow ^7F_4$	710	0.53	-0.52	0.45	0.46

Table 3 Overall europium-centred luminescence quantum yields and sensitisation efficiencies for [Eu₂L₃] helicate complexes recorded in CH₃CN.²⁴

Eu ₂ L ₃	τ_{obs} , ms	Φ_{tot} , %	τ_{R} , ms	$\Phi_{\text{Ln}}^{\text{Ln}}$, %	η_{sens} , %
L = 1(S,S)	1.44 ± 0.01	5.8 ± 0.1	6.9 ± 0.4	21.0 ± 0.5	27.6
L = 4(R,R)	1.48 ± 0.01	3.2 ± 0.5	10.4 ± 0.7	14.3 ± 0.8	22.4
L = 5(S,S)	1.43 ± 0.03	2.6 ± 0.1	9.4 ± 0.8	15.3 ± 0.9	17.0

η_{sens} – antenna-to-ion energy transfer efficiencies, τ_{obs} – the observed luminescence lifetime, Φ_{tot} – the experimental overall luminescence quantum yield ($\lambda = 279$ nm), τ – the radiative lifetime, Φ_{Ln} – the ex R Ln intrinsic quantum yield of lanthanide.

Photophysical properties of [Eu₂:L₃] assemblies in solution

The photophysical properties of the triple stranded di-metallic helicates [Eu₂:L₃], were monitored in solution upon the self- assembly formation in real time. These were seen to be close to identical for each pairs of enantiomers, therefore we report herein, only the photophysical results for one of each enantiomers (Table 3 and Fig. S32[†]). For these complexes, the life- times were recorded in solution by monitoring the decay of the Eu(III) 5D₀ → 7F₂ band upon ligand excitation at 279 nm. In all the cases, the lifetime decay curves were best fitted to mono- exponential decay using Origin9.1®; the values being practically the same, ca. 1.45 ms, for all the complexes (Scheme 1), which also corresponds to the life- time values obtained for the addition of 0.67 equivalents of metal ion to the solution. This

suggest the presence of one type of coordination environment around Eu(III) ion in solution (Table 3 and Fig. S33†). The total quantum yield (Φ_{tot}) of the Eu(III) emission was also determined for these self-assembly species. The results showed that Φ_{tot} was two times higher for $[\text{Eu}_2(1(\text{S},\text{S}))_3]$ than $[\text{Eu}_2(4(\text{R},\text{R}))_3]$ and $[\text{Eu}_2(5(\text{S},\text{S}))_3]$. The values of the radiative lifetime (τ_{R}), the intrinsic quantum yield (Φ_{Ln}) and the antenna-to-ion energy Ln transfer efficiencies (η_{sens}) calculated for Eu(III) assemblies using their emission spectra, the observed luminescence lifetime (τ_{obs}), and Φ_{tot} suggests that non-radiative processes was most apparent for $L = 1(\text{S},\text{S})$. However, as the same system reveals the highest Φ_{Ln} and η in comparison to the other Ln sens two systems means that the radiative processes in this case compete well with non-radiative ones. This can be due to the better sensitisation efficiency thanks to the ligand structure allowing the position of triplet state level relative to Eu(III) levels as well as better geometrical arrangement of the ligand around the metal ion.^{1,25}

Mass-spectrometry analysis of $[\text{Eu}_2:\text{L}_3]$ assemblies

Mass-spectrometry analyses were performed using $[\text{Eu}_2:\text{L}_3]$ complexes formed under thermodynamic control by reacting 1 equiv. of the respective L with 0.67 equiv. of $\text{Eu}(\text{ClO}_4)_3$ (50 wt% in water) in CH_3OH at 65 °C under microwave irradiation for 2 hours. The complexes were precipitated out of solution using ether diffusion, dried under vacuum before being subjected to matrix-assisted laser desorption/ionisation mass-spectrometry (MALDI-MS) with DCTB used as matrix (Fig. 6 and S34–S37†). For the Eu(III) samples investigated in this work the presence of $[\text{Eu}:\text{L}]$, $[\text{Eu}_2:\text{L}]$, $[\text{Eu}:\text{L}_2]$, $[\text{Eu}_2:\text{L}_2]$ species was observed. At the same time $[\text{Eu}_2:\text{L}_3]$ species were only observed for the samples with ligands $L = 3(\text{S},\text{S})$ and $4(\text{R},\text{R})$. The presence of tri-metallic species was also observed for complexes with $L = 3(\text{S},\text{S})$ $[\text{Eu}_3:\text{L}_3]$. The samples with $L = 6(\text{R},\text{R})$ also resulted in the assignment of a $[\text{Eu}_3:\text{L}_2]$ species, suggesting possible formation of more complex systems, but such higher

order systems have been reported in the literature by others,¹² believed to be due to supramolecular cooperativity effects.^{23b,26} Unfortunately, we were unable to get reliable MS or HRMS results when using samples from the various spectroscopic titrations discussed above, this most likely being due to low concentrations of these assemblies in these samples.

¹H NMR study of the self-assembly between L and Eu(III)

We also investigated the formation of these helicates in solution by observing the changes in the ¹H NMR of the ligands. Unlike commonly seen for many other lanthanide complexes (such as cyclen based macrocyclic complexes), the changes in the NMR of the systems studied here upon addition of the lanthanides, occurred over relatively narrow range, and with less signal broadening, allowing for the formation of these assemblies to be monitored in real-time. The solutions of 4 (R,R) and 6(R,R) ($c = 1.00 \times 10^{-3}$ M) were titrated against gradual addition of Eu(CF₃SO₃)₃ (0 → 1.99 or 0 → 1.11 equiv., correspondingly) in CD₃CN. In case of 4(R,R) (Fig. 7 and S38†), noticeable changes started to occur upon addition of 0.22 equiv. of Eu(III), and these continued until the addition of 0.66 equiv., after which no significant changes were observed within the aromatic region. This confirming the successful formation of the [Eu₂:43] in solution at these concentrations. Furthermore, we also observed that a resonance assigned to water, that was present in the solution of 4(R,R) at 2.13 ppm, experienced a significant upfield shift to 1.74 ppm (Fig. S39†) suggesting a possible encapsulation of water within the helicate. As we anticipate a full saturation around the two lanthanide ions (by the three ligand strands), it is possible that the water molecule is located within a cavity, that is generated at the centre of the structure where the three bridges are located. However, in the absence of a crystal structure of this [Eu₂:L₃] assembly, this cannot be fully determined. The ¹H NMR spectrum was recorded for Eu(III) complex with L = 3(S,S) synthesised under thermodynamic control and its shape closely resembles the

spectrum observed for 0.66 equiv. addition of Eu(III) ion to the ligand during the titration (Fig. S40†) therefore confirming the formation of [Eu₂:L₃] in both cases together with MALDI-MS data. In contrast to these results, then upon addition of Eu(CF₃SO₃)₃ to 6(R,R) significant broadening was observed for most of the characteristic signals in the ¹H NMR spectrum beyond the addition of 0.44 equiv. of Eu(III), making it impossible to distinguish between these signals (Fig. S41†). However, similarly to the aforementioned titration it was possible to follow the changes of the water peak where it started to shift upon addition of 0.66 equiv. of Eu(III) due to the possible encapsulation within [Eu_n:L_m] assemblies (Fig. S42†). The broadening of the ¹H NMR spectra during the self-assembly process between 6(R,R) and Eu(III) occurs most likely due to the similar exchange rate between metal and ligand during complexation compared to the NMR time scale, which could be directly related to the size of the central spacer unit.²⁷ The signals in the ¹H NMR spectrum of Eu(III) complex with L = 5(S,S) made under thermodynamic control revealed broad signals similarly to the titration data and along with the exchange rate this can occur due to the steric hindrance of the ligand during the assembly processes (Fig. S43†). However, as the solution of the complex was heated up to 40 °C and then cooled down to 22 °C it became more resolved (Fig. S44†) as heating increased the reaction rate shifting the equilibrium from L and Eu(III) towards the formation of highly symmetrical [Eu₂:L₃] (L = 5(S,S)) species.²⁷

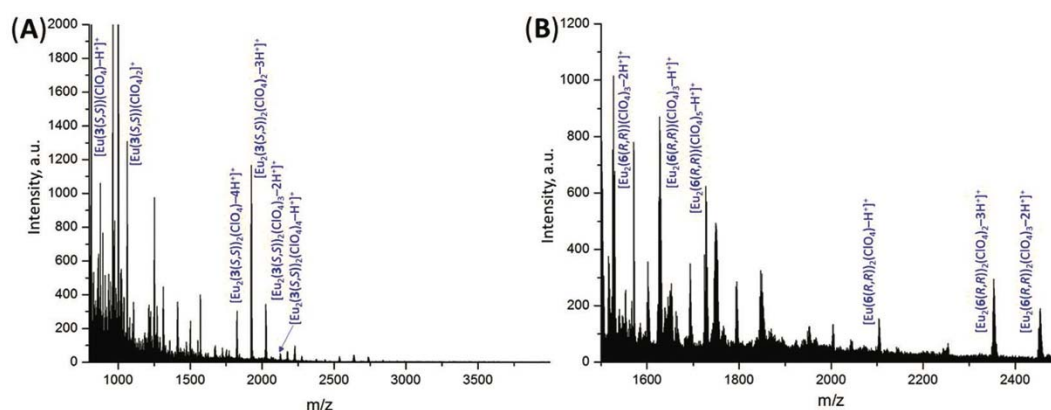


Fig. 6 MALDI-MS spectra of Eu(III) complexes with (A) L = 3(S,S), (B) L = 6(R,R).

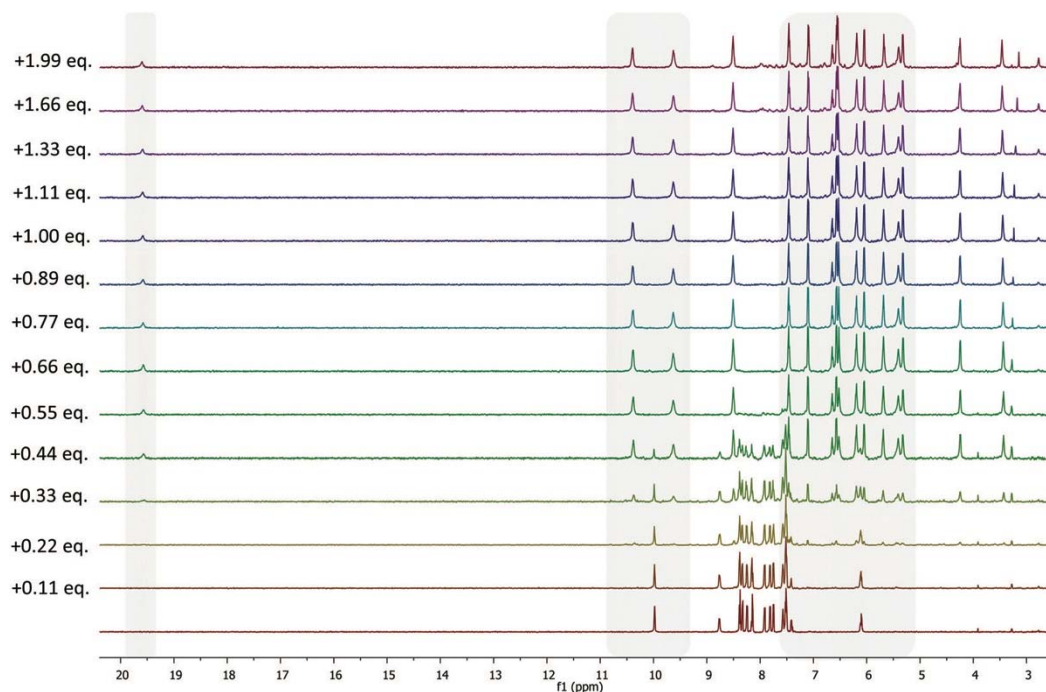


Fig. 7 The changes in ^1H NMR of 4(R,R) ($c = 1.00 \times 10^{-3}$ M) upon addition of $\text{Eu}(\text{CF}_3\text{SO}_3)_3$ ($0 \rightarrow 1.99$ equiv.) in CD_3CN (600 MHz, $T = 25^\circ\text{C}$).

Conclusions

In conclusion we have synthesised four new ligands by coupling the diamine linker of various size with two chiral half-helicate ligands. The titrations of these at low concentration upon addition of Eu(III) in CH_3CN allowed us to determine the binding constants values for $[\text{Eu}:\text{L}]$, $[\text{Eu}_2:\text{L}_2]$ and $[\text{Eu}_2:\text{L}_3]$ self-assembly species. Lower values of the binding constants were found for the systems with the bulkier spacer 5(S,S), 6(R,R) as the resulting structures were most likely less rigid in comparison to 3(S,S), 4(R,R). The conformational changes of these chiral ligands upon complexation with Eu(III) ions were monitored by following the

changes in the CD spectra. Recording CPL spectra for [Eu₂:L₃] allowed us to assign absolute configuration of Eu(III) centres and calculate the dissymmetry factor values.

The luminescence quantum yields between [Eu₂:L₃] with L = 4(R,R) and 5(S,S) assemblies were surprisingly similar with the highest quantum yield being observed for Eu(III) complex with 1,3-phenylenedimethanamine linker (L = 1(S,S)).

The mass-spectrometry analysis of the solid complexes shown the presence of [Eu₂:L₃] assemblies. However, the results revealed the presence of [Eu₃:L₂] and [Eu₃:L₃] assemblies suggesting the presence of more complicated species in the solid state at high concentration. ¹H NMR titration study performed at higher concentration than spectroscopical studies confirmed the formation of highly symmetrical [Eu₂:L₃] helical species in solution for the systems with L = 3(S,S), 4(R,R) while indicating slower kinetics of the complexes formation for the ligands with bulkier spacer L = 5(S,S), 6(R,R).

In summary, the formation of chiral, luminescent [Eu₂:L₃] assemblies with high stability constants has been demonstrated in solution, while the composition of the assemblies in the solid state was more complex. The diversity of these systems allowed us to investigate their formation using various techniques and help their development towards materials for the applications in electronics for virtual reality applications (3D screens) and in medical diagnostic. We are exploring these avenues in greater details.

Conflicts of interest

There are no conflicts of interest to declare.

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